

Heat of Solution

Objective:

When certain compounds are dissolved in water there is an enthalpy change (heat change) associated with this process. In this experiment you will determine the heat of solution of compounds like potassium nitrate and ammonium chloride. Compare their results and come up with logical arguments as to the reason for the observation. This experiment can be performed using improvised calorimeters (Styrofoam cups)

- Research and come up with a possible lab procedure for setting up your equipments and performing the experiment.
- Include safety concerns that should be addressed
- Make a list of equipments that may be used to perform the experiment.
- Submit a draft of the procedure adopted for the lab to your teacher and discuss if the materials required are available and the lab activity can be performed.
- Choose a convenient time for your experiment, record your data and present your findings as instructed.